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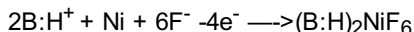
Electrochemical fluorination of organic compounds.

3. Anode process nature in ECF.

In the first publications devoted to ECF, Simons[9] has shown that the process takes place without elemental fluorine evolution. The fluorine evolution begins only at an electrode voltage above 7-8v, that leads to the full destruction of the organic substrate to be fluorinated and accompanied with explosions in the electrolyzer. Nevertheless authors [32] postulate that the process takes place by means of fluorine ion discharge on the anode and secondary interactions of the fluorine radical with an organic molecule. Similar ideas were developed by Okatov [40] Knunians and Sokolsky [41] who assumed even fluorine radical diffusion into the electrolyte mass.

To explain an induction period prior to the beginning of organofluorine products evolution, Haszeldine et al. [42] in their study of the electrochemical fluorination of sulfur-containing organic compounds assumed that the fluorination was carried out through the forming nickel fluorides on the anode. Then the mechanism of ECF is reduced to electron transfer by a higher metal fluoride similar to the fluorination by cobalt trifluoride [4]. But special investigations of behavior of different anode materials in ECF [43] have shown that this process does not occur on cobalt and (being only of anode specific) takes place only on nickel and to a lesser degree on its alloys. The nickel electron structure suggests the presence of nickel dioxide in alkali medium, which is the basis of alkali batteries, but compounds of oxidized nickel in an acid medium are not known [2]. But when a mixture of nickel chlorides and potassium chlorides was treated with fluorine at heating, crystal structures of hexafluoronickelates of heavy alkali metals were produced and studied[44].

Since the organic compounds exist in AHF medium as protonated cations, Burdon and Tatlow [45] and later others [46,47] assumed that the ECF process occurs with participation of complex higher nickel fluorides in which the nickel of the highest valence acts as an anion complexing center, fluorine atoms act as ligands and the protonated organic compound acts as a cation according to the scheme:



One can assume a possibility of intramolecular transfer of electrons in this complex and, as a result of this, nickel passes into 2-valent state and fluorine replaces hydrogen in the organic compound with HF evolution.

Rozhkov [11] and Schmidt [2] considered the mechanism of carbocation formation in electrochemical oxidation of an organic molecule and an addition of fluorine anion proved for the fluorination on Pt anode in a medium of acetonitrile and acetic acid. They made an attempt to generalize this mechanism with that of electrochemical fluorination in AHF medium. But it seems not productive to prove the unified mechanism in such different media and on anodes of different structure.

As it was mentioned, the simplest process of electrochemical fluorination is water fluorination with formation of fluorine monoxide. Hygroscopicity of AHF and moisture condensation on a low temperature reflux condenser of the electrolyzer signify that moisture is almost always present at the beginning of ECF. During the electrolysis water is fluorinating and according to author [9] its content in the electrolyte after some hours does not exceed 0.1%, while the effectiveness of the process comes down significantly when water content increases to 1% and at a value of 10% there is the formation of explosive mixtures in the electrolyzer. According to Simons [48], it is possible to produce perfluorooctane in 30% yield using water conductive properties in the octane electrolysis in the presence of 1.5% of water.

In this connection the mechanism of water oxidation in AHF and water influence on the ECF process have been a subject of researchers' steady attention. In particular Engelbrecht [49] studied the water electrolysis within a wide range of concentrations and has shown that at a moisture content above 30%, the only anodic process is water discharge with formation of oxygen and ozone. As the concentration of water decreases, the oxydation and fluorination start to compete and at a concentration below 20% there is found up to 60% of fluorine monoxide. Rogers et al. [50] studied hydrogen fluoride drying by electrolysis and have shown that the current drop in the electrolysis of moist hydrogen fluoride is determined not only by the moisture content reduction but also by a change of the anode condition during the electrolysis, so full drying can be achieved only using a reverse current. Nagase et al. [51] studied water effect on the process of electrochemical fluorination of alcohols and esters of carboxylic acids. It has been determined that the presence of moisture in amount of 0.2-0.7% in the electrolyte affects negligibly the perfluoroacyl fluoride yield, but the current yield of these products decreases significantly though moisture removal was made earlier than fluorination of the main substrate.

Different authors tried to prove their assumptions of the fluorination mechanism studying the content of the fluorination products but strong arguments of the nature and mechanism of the electrochemical process could be obtained only using modern electrochemical methods of investigation of electrode processes.

Their realization has been limited for a long time by a possibility to measure the anodic potential in AHF versus a stable nonpolarizable reference electrode.

For the first time Fredenhagen [52] measured in AHF the electrode potential difference of gas electrodes (hydrogen and fluorine) on platinum and determined experimentally the standard fluorine potential equal to 2.76V at 0°C.

Koerber and De Vries [53] have measured the equilibrium electrode potentials of metal fluoride electrodes in AHF versus mercurous fluoride electrode, they have determined EMF for chains without migration of the following types: $M(Hg)|MF_2(s)||HF(NaF)||Hg_2F_2||Hg$ where $M=Cd; Cu$, and also $Pb(Hg)|PbF_2 \cdot 2.5HF(s)||HF(NaF)||Hg_2F_2||Hg$ and for chains with migration of $Hg|Hg_2F_2||HF(NaF)||HF(NaF;H_3OF;AgF)||Ag$ type. The authors showed that in AHF medium it is possible to determine emf with an accuracy of 0.1 mV and the concentration dependencies of the electrode potentials are in accordance with Nernst's equation. It was shown that lead fluoride in this medium formed an insoluble complex, also free energies of formation of metal fluorides and the values of the activity coefficients of silver fluoride solutions in AHF were determined. Some attempts [54,55] were made on a polarographic investigation of ECF but they were of little information.

Hackerman et al. [56] first have used a mercurous fluoride reference electrode in polarization measurements in AHF medium whereas Kaurova [59] and Watanabe [57,58] used a hydrogen reference electrode. The apparatus illustrated in **Fig.3** is that designed by Watanabe, the cell of Kaurova's design is illustrated in **Fig.4**. Later Burrows [60] and Novak [61] developed for this purpose a cupric fluoride electrode and showed that the latter is the most convenient and stable reference electrode in this medium.

Using water as a fluorination model, Hackerman and co-workers made potentiostatic and halvanostatic investigations of anodic processes and measured the differential capacity of the double layer in the anhydrous and moist hydrogen fluoride. They showed with evidence that metal dissolution with formation of a passive film of nickel fluoride takes place on nickel in the anodic range of the potentials. When the potential reaches a value of fluorine evolution, the anode current, according to Tafel's equation at a slant b equal to 600mV, conforms to the only process within the range of high anodic potentials which is the process of fluoroanion discharge with formation of fluorine radical which being adsorbed in the nickel fluoride film, begins to form molecules with elemental fluorine evolution only at potentials above 5.5V. These ideas were confirmed both by electrochemical measurements in AHF in the presence of moisture [57,59] and by the study of fluorine absorption on nickel fluoride [58]. There was determined the energy of fluorine absorption on nickel fluoride within the range of temperatures of (-72 +27°C), its value (below 4 kJ/mol) together with the slope of Tafel's curve are an evidence that the limiting stage of the process is fluorine ion discharge. This research has confirmed the extremely high value of b and hence a value of the electron transfer coefficient α calculated according to the equation for b : $b = 2.3 RT/(1 - \alpha)zF$, which varies from 0.8 to 0.9 (according to different authors) and points to an energy barrier distortion. This phenomenon is typical enough for the processes in the range of high anodic potentials associated with radical absorption on the electrode surface covered with a film of oxides as in Kolbe's synthesis [62].

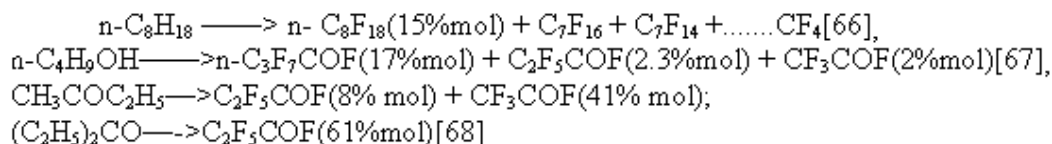
Serushkin and co-workers [63] studied charging curves on the nickel anode in AHF in the presence of moisture and acetic acid. Considering the electricity amount corresponding to the delays on the curves, they determined that the filling of the nickel surface is equal to $1 \cdot 10^{-8}$ mol/cm² which is in conformity with formation of 3-4 layers of nickel fluoride on the anode. But a difference between the charging curves in solutions of acetic acid and water do not point to a change of the anode process nature.

In the presence of moisture, anodic nickel dissolution comes down significantly, the current yield of the corrosion process and passivation potential decrease, all this is associated by the authors with complexation of nickel fluoride with water (Clifford [64]). In dependence on the moisture content, nickel fluoride exists in this medium as aquacomplexes: dehydrate, tetrahydrate and hexahydrate. Water which is protonated in this medium and exists as a hydroxonium cation, at interaction with nickel fluoride is subjected to deprotonation. Earlier there was found complexation of various organic compounds with metal fluorides [65], especially with nickel fluoride which complexation with acetonitrile was studied quantitatively and there was found the formation of insoluble solid phase of $Ni(CH_3CN)_nF_2$ at increasing the acetonitrile concentration up to 1M.

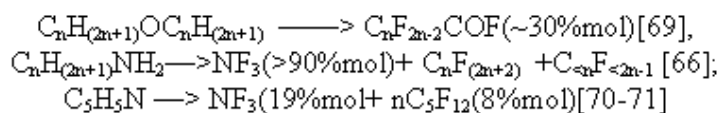
Complexation of nickel fluoride with organic compounds accompanying their deprotonation on the anode surface obviously removes the electric barrier preventing cation participation in the anodic reaction at high potentials.

It followed from the first Simons' publication [66] that a process of ECF is not selective in most cases. It is accompanied by the following processes:

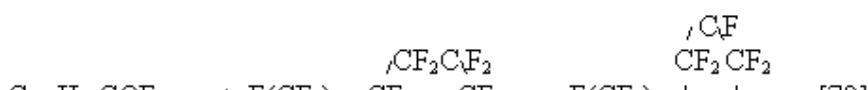
1. destruction of the carbon skeleton of the starting compound:

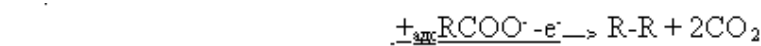


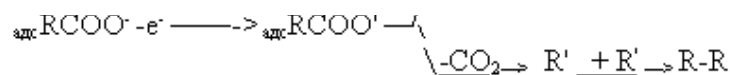
2. bond cleavage of carbon with a heteroatom:



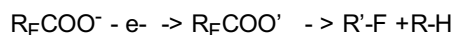
3. cyclization:





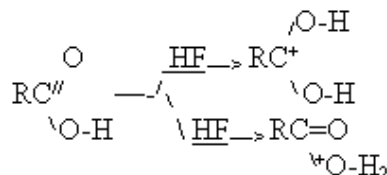


The following scheme was proposed [45]:



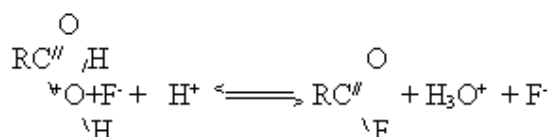
which is open to criticism because there are no anions of carboxylic acids in AHF medium and perfluorinated structures are adsorbed from the anode surface. A distinguishing feature of Kolbe's electrolysis is radical dimerization slightly observed at electrochemical fluorination and it is difficult to consider hydrogen fluoride disproportionation under the action of radicals.

A source of unsatisfactory yields of perfluoroacylfluorides in the ECF of carboxylic acids one should search in the structure of the acid cation complex which is formed at acid protonation. Obviously, two proton-deficient neighboring centers with lone-pair electrons exist in an acid molecule, therefore two structures are possible for the protonated complex. In the first of them, a positive charge is concentrated on the carbon atom in the structure of the carbocation; in the other structure it is concentrated on the oxygen atom in the structure of oxonium ion:

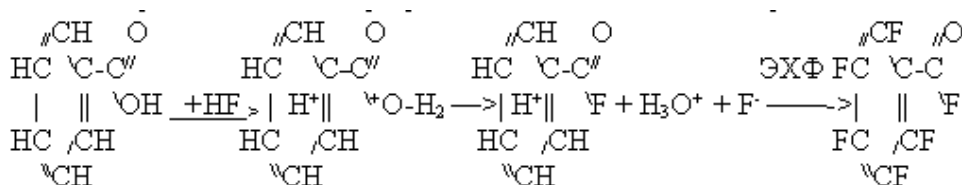


Deprotonation of the carbon atom with adsorption of the acid molecule on nickel fluoride at fluorination will result in the formation of structure -C-O-F which is stabilized with fluorine monoxide formation and carbonyl elimination in the form of CO or COF_2 .

Deprotonation of the oxygen atom may be accompanied with solvolysis with formation of appropriate acylfluoride and hydroxonium cation followed by subsequent acylfluoride fluorination:



If there is a second proton-deficient centre in the organic radical, for example an aromatic ring or alkoxy group, repulsion of similar charges makes more preferable the formation of oxonium ion whose positive charge is by a distance of C-O bond length further away than carbocation. In this case acylfluoride formation becomes more easy and as a result yields of perfluoroacylfluorides in fluorination of benzoic acid [97,98] and alkoxycarbonic acids [92] are higher than those for typical carboxylic acids:



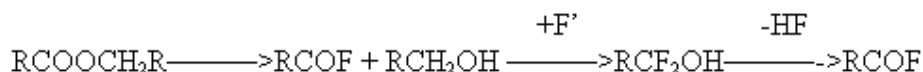
Electrochemical fluorination of esters of carboxylic acids was succeeded in relatively high yields of perfluoroacylfluorides. Thus, methyl butyrate fluorination [99] gave perfluorobutyl fluoride in 35% mol yield, which is a little different in comparison with the results of butyryl fluoride fluorination [91] and is in good agreement with the data of S. Nagase et al. [100] given in the table

Table

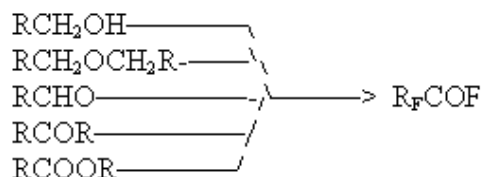
Starting substrate	Yield, %mol		
	CF_3COF	$\text{CF}_3\text{CF}_2\text{COF}$	$\text{CF}_3\text{CF}_2\text{CF}_2\text{COF}$
$\text{CH}_3\text{COOCH}_2\text{CH}_3$	52.3x2	-	-
$\text{CH}_3\text{COOCH}_2\text{CH}_2\text{CH}_3$	19	24	-
$\text{CH}_3\text{COOCH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	22	1.0	15
$\text{CH}_3\text{CH}_2\text{COOCH}_2\text{CH}_3$	20	23	-
$\text{CH}_3\text{CH}_2\text{COOCH}_2\text{CH}_2\text{CH}_3$	2.1	3362	-
$\text{CH}_3\text{CH}_2\text{COOCH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	2.4	25	12
$\text{CH}_3\text{CH}_2\text{CH}_2\text{COOCH}_2\text{CH}_3$	18	1.0	16
$\text{CH}_3\text{CH}_2\text{CH}_2\text{COOCH}_2\text{CH}_2\text{CH}_3$	1.7	16	11
$\text{CH}_3\text{CH}_2\text{CH}_2\text{COOCH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	2.3	2.0	28,5x2

The process of ester fluorination includes ester solvolysis with formation of acylfluoride and an alcohol which after hydrogen

The process of ester fluorination includes ester solvolysis with formation of acylfluoride and an alcohol which after hydrogen substitution with fluorine at α -carbon atom undergoes dehydrofluorination with formation of acylfluoride group according to the scheme:



Simons and co-workers in fluorination of various oxygen-containing compounds such as alcohols [39], ethers [102], ketones [66] identified only fluorocarbons in the electrolysis products. Later Nagase and co-workers after broad investigations of ECF of oxygen-containing compounds showed that all of them form perfluoroacylfluorides as a result of electrolysis [67-69, 100, 101].

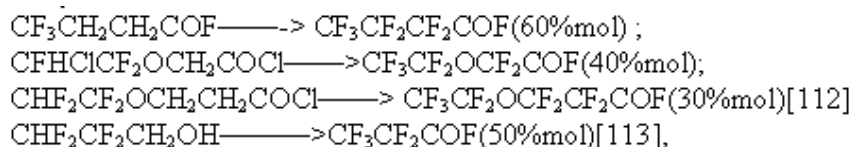


The dependence of perfluoroacylfluoride substance yield on the carbon chain length and the functional group is shown in **Fig.6**.

We mentioned above the differences in yields of perfluorinated compounds in dependence on their structure, when primary, secondary amines and pyridine are fluorinated with rupture of C-N bonds, whereas from aliphatic tertiary amines [102], N,N - dialkylanilines and cyclohexylamines [71, 106, 107], salts of quaternary N-alkylammonia heterocyclic bases [108, 109] there are produced perfluoro analogs with high yields. The dependence of perfluoroanalog yield on the chain length of N-alkyl substituent according to data [106] is given in **Fig.7**.

Before we mentioned producing perfluoroalkansulfofluorides from conforming alkansulfohalogenides, but it is not complete data on ECF of sulfur-containing compounds. Electrolysis of solutions of dimethylsulfide and carbon disulfide in AHF results in the formation mainly of sulfur trifluoromethylpentafluoride in 57 and 90%mol yield [110]. A yield of sulfur bistrifluoromethyltetrafluoride did not exceed 5%mol. Further by ECF of dialkylsulfides and -disulfides, also cyclic thioethers [111] there were produced appropriate sulfur perfluoroalkylpentafluorides and sulfur bisperfluoroalkyltetrafluorides in good yields, moreover the yield of sulfur bisperfluoroalkyltetrafluorides increases up to that of pentafluorides with the chain growth.

Almost in all cases of ECF of partially fluorinated compounds, yields of the electrolysis products increase considerably. We mentioned the fluorination of difluorocyclohexane [80] when the yield of perfluoro-C6 structures is noticeably higher than for hexane. Similarly:



Using this advantage of partially fluorinated compounds, perfluoroenanthic and perfluoropelargonic acids by ECF of telomer alcohols of $\text{H}(\text{CF}_2\text{CF}_2)_n\text{CH}_2\text{OH}$ formula where $n=3,4$ are produced in Russia.

(to be continued)